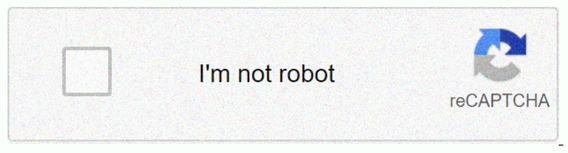

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Lab Safety 1

You will usually find a set of 'Lab Rules' in a science lab. They are there for your safety!

The drawing below shows a lab where there are no safety rules.

1 Write down a list of all the things going wrong in this lab.
2 For each of your answers, write a safety rule.
3 Design a poster for your lab, showing one safety rule and the reason for having it.

[Developing skills in: planning]

Name Ms. Crittenden ANSWER KEY
Date _____

pH and pOH Worksheet

1. Calculate the pH and pOH of solutions having the following concentration:
   a. 0.00010 mole H_3O^+ per liter. pH = 4 pOH = 10
   b. 0.010 mole OH^- per liter. pH = 12 pOH = 2
   c. 1.0×10^{-5} mole OH^- per liter. pH = 9 pOH = 5
   d. 1.0×10^{-2} mole H_3O^+ per liter. pH = 2 pOH = 12

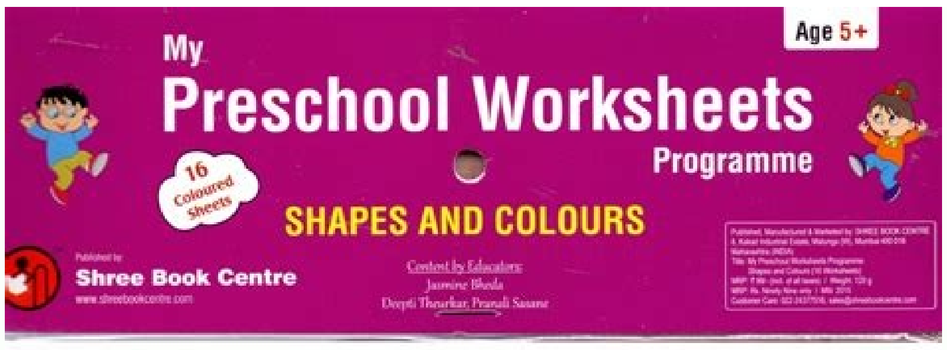
2. Calculate the $[H_3O^+]$ of the following solutions:
   a. pH = 3.0 $[H_3O^+] = 1.0 \times 10^{-3}$ M
   b. pH = 6.0 $[H_3O^+] = 1.0 \times 10^{-6}$ M
   c. pOH = 12.0 → pH = 2.0 $[H_3O^+] = 1.0 \times 10^{-2}$ M

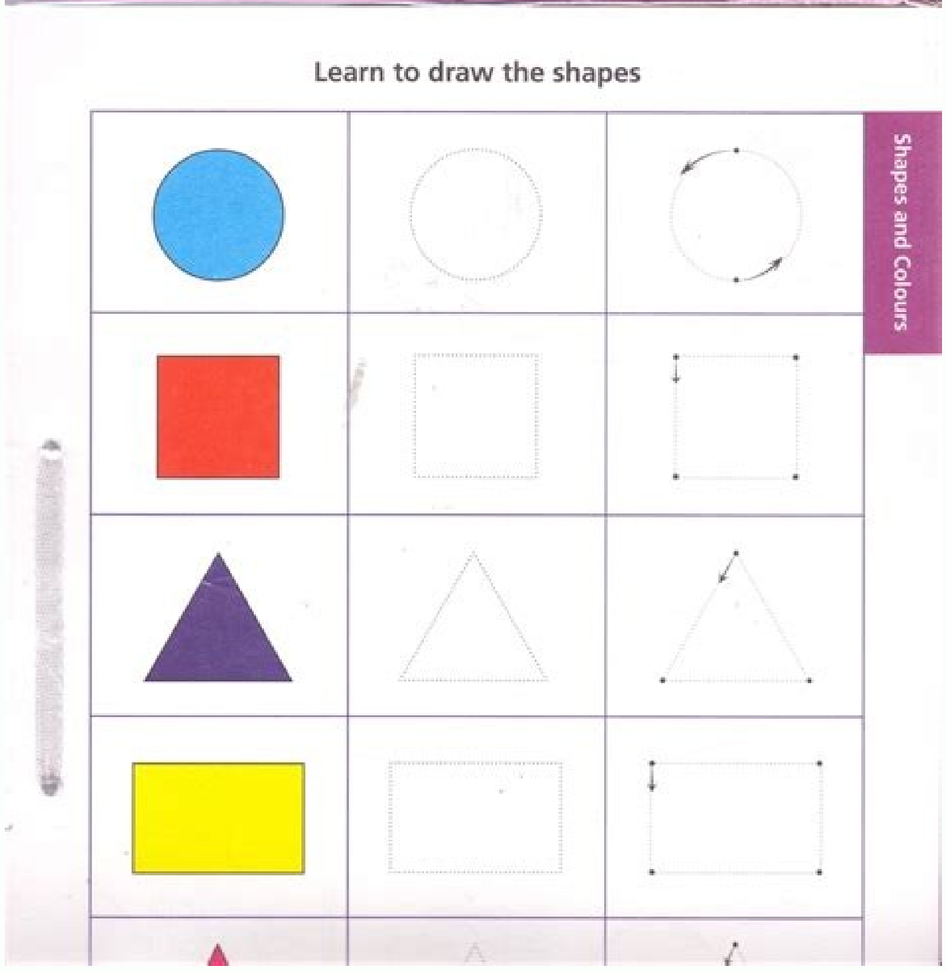
3. Calculate the $[OH^-]$ of the following solutions:
   a. pOH = 11.0 $[OH^-] = 1.0 \times 10^{-11}$ M
   b. pH = 4.0 → pOH = 10.0 $[OH^-] = 1.0 \times 10^{-10}$ M
   c. pOH = 8.0 $[OH^-] = 1.0 \times 10^{-8}$ M

4. Calculate the pH and pOH of solutions having the following concentraticis. Assume 100% ionization. Remember that 1 mole of H_2SO_4 produces 2 moles of H_3O^+ ion.
   a. 0.0025 M NaOH pH = 11.4 pOH = 2.6
   b. 0.0025 M H_2SO_4 ×2 pH = 2.3 pOH = 11.7
   c. 0.075 M H_2SO_4 ×2 pH = 0.82 pOH = 13.18
   d. 0.048 M HCl pH = 1.3 pOH = 12.7
   e. 0.032 M KOH pH = 12.5 pOH = 1.5
   f. 0.00017 M NaOH pH = 10.2 pOH = 3.8

5. Calculate the $[H_3O^+]$ and $[OH^-]$ of the following solutions:
   a. pH = 2.500 → pOH = 11.5 $[H_3O^+] = 0.0032$ M $[OH^-] = 3.2 \times 10^{-12}$ M
   b. pOH = 5.800 → pH = 8.2 $[H_3O^+] = 6.3 \times 10^{-9}$ M $[OH^-] = 1.6 \times 10^{-6}$ M
   c. pOH = 3.200 → pH = 10.8 $[H_3O^+] = 1.6 \times 10^{-11}$ M $[OH^-] = 6.3 \times 10^{-4}$ M
   d. pH = 4.700 → pOH = 9.3 $[H_3O^+] = 2.0 \times 10^{-5}$ M $[OH^-] = 5.0 \times 10^{-10}$ M
   e. pH = 9.600 → pOH = 4.4 $[H_3O^+] = 2.5 \times 10^{-10}$ M $[OH^-] = 4.0 \times 10^{-5}$ M
   f. pOH = 10.300 → pH = 3.7 $[H_3O^+] = 2.0 \times 10^{-4}$ M $[OH^-] = 5.0 \times 10^{-11}$ M

6. The approximate pH of some common substances is lised. Calculate the pOH, the $[H_3O^+]$ and the $[OH^-]$
   a. Vienger 2.8 14 - 2.8 pOH = 11.2 $[H_3O+] = 0.0016$ M $[OH^-] = 6.3 \times 10^{-12}$ M
   b. Orange 3.5 14 - 3.5 pOH = 10.5 $[H_3O^+] = 3.2 \times 10^{-4}$ M $[OH^-] = 3.2 \times 10^{-11}$ M
   c. Rainwater 6.2 14 - 6.2 pOH = 7.8 $[H_3O^+] = 6.3 \times 10^{-7}$ M $[OH^-] = 1.6 \times 10^{-8}$ M
   d. Seawater 8.5 14 - 8.5 pOH = 5.5 $[H_3O^+] = 3.2 \times 10^{-6}$ M $[OH^-] = 3.2 \times 10^{-9}$ M
   e. Soft drink 3.0 14 - 3.0 pOH = 11.0 $[H_3O^+] = 0.001$ M $[OH^-] = 1.0 \times 10^{-11}$ M
   f. Tomato 4.2 14 - 4.2 pOH = 9.8 $[H_3O^+] = 6.3 \times 10^{-5}$ M $[OH^-] = 1.6 \times 10^{-10}$ M
   g. Egg 7.8 14 - 7.8 pOH = 6.2 $[H_3O^+] = 1.6 \times 10^{-7}$ M $[OH^-] = 6.3 \times 10^{-8}$ M
   f. Milk of magnesia 10.5 14 - 10.5 pOH = 3.5 $[H_3O^+] = 3.2 \times 10^{-11}$ M $[OH^-] = 3.2 \times 10^{-4}$ M





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Enclosed here is a brief description and a circuit diagram to illustrate each type of circuit. How fun, and maybe a little terrifying, that would be! 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